Name……………………...........

Form……………………………

Naming compounds and balancing equations.

1. Naming compounds. Fill in the gaps in the table below.

An example has been done for you.

|  |  |  |
| --- | --- | --- |
| **Element** | **Element** | **Compound** |
| Copper | Chlorine | Copper chloride |
| Magnesium | Chlorine |  |
| Aluminium | Oxygen |  |
| Lead | Iodine |  |
|  |  | Zinc oxide |
|  |  | Calcium nitride |

 (7 marks)

1. Naming compounds part 2. Fill in the gaps.

|  |  |  |  |
| --- | --- | --- | --- |
| **Element** | **Element** | **Element** | **Compound** |
| Copper | Sulfur | Oxygen | Copper sulfate |
| Silver | Nitrogen | Oxygen |  |
| Copper | Carbon | Oxygen |  |
|  |  |  | Lead nitrate |
|  |  |  | Potassium chlorate |

 (8 marks)

1. Atoms in compounds. Fill in the gaps.

|  |  |  |
| --- | --- | --- |
| **Compound** | **Formula** | **Atoms** |
| Carbon dioxide | CO2 | 1 carbon atom2 oxygen atoms |
| Water |  |  |
|  | CuO |  |
|  | Al2O3 |  |
|  | Ca(NO3)2 |  |

 (8 marks)

1. Balancing equations. Write a balanced symbol equation for the following reactions:
2. Magnesium + oxygen 🡪 magnesium oxide

…………………………………………………………………………………….

1. Copper + oxygen 🡪 copper oxide

…………………………………………………………………………………….

1. Sodium + oxygen 🡪 sodium oxide

…………………………………………………………………………………….

1. Iron + oxygen 🡪 iron (III) oxide

…………………………………………………………………………………….

 (8 marks)

Extension challenge – put the correct numbers into these to balance the equations.

1. …. H2(g) + …. Cl2(g) à …. HCl(g)
2. …. H2(g) + …. O2(g) à …. H2O(l)
3. …. N2(g) + …. H2(g) à …. NH3(g)
4. …. SO2(g) + …. O2(g) à …. SO3(g)
5. …. ZnS(s) + …. O2(g)  à …. ZnO(s)  + …. SO2(g)
6. …. Fe2O3(s) + …. Al(s) à …. Al2O3(s) + …. Fe(s)